



Cambridge Assessment International Education

Cambridge International General Certificate of Secondary Education

CANDIDATE NAME			
CENTRE NUMBER		CANDIDATE NUMBER	
CHEMISTRY		0620	/61
Paper 6 Alterna	ative to Practical	October/November 2	019
		1 h	our
Candidates ans	swer on the Question Paper.		

READ THESE INSTRUCTIONS FIRST

No Additional Materials are required.

Write your centre number, candidate number and name on all the work you hand in.

Write in dark blue or black pen.

You may use an HB pencil for any diagrams or graphs.

Do not use staples, paper clips, glue or correction fluid.

DO **NOT** WRITE IN ANY BARCODES.

Answer all questions.

Electronic calculators may be used.

You may lose marks if you do not show your working or if you do not use appropriate units.

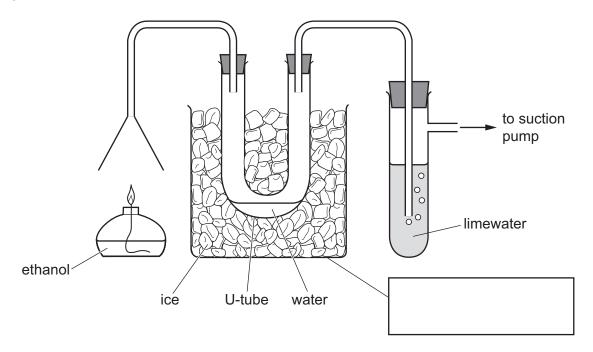
At the end of the examination, fasten all your work securely together.

The number of marks is given in brackets [] at the end of each question or part question.

This syllabus is regulated for use in England, Wales and Northern Ireland as a Cambridge International Level 1/Level 2 Certificate.



The apparatus shown was used to investigate the products formed when ethanol, $\mathrm{C_2H_5OH}$, burns 1 in air.



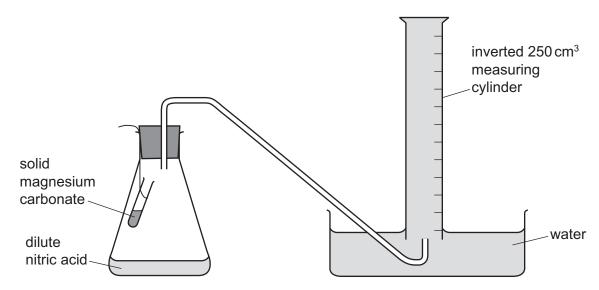
(a)	Complete the box to name the apparatus.	[1]
(b)	Explain why water collects in the U-tube.	
(c)	State a chemical test for water.	
(0)	test	
	observations	 [2]
(d)	What is the purpose of the suction pump?	
		ניו
(e)	Predict the change seen in the test-tube containing the limewater. Explain your prediction.	
		[2]

[Total: 8]

2 A student investigated the rate of reaction between dilute nitric acid and an excess of solid magnesium carbonate at room temperature.

The apparatus was set up as shown in the diagram.

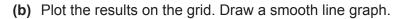
A small test-tube containing magnesium carbonate was suspended in the conical flask.

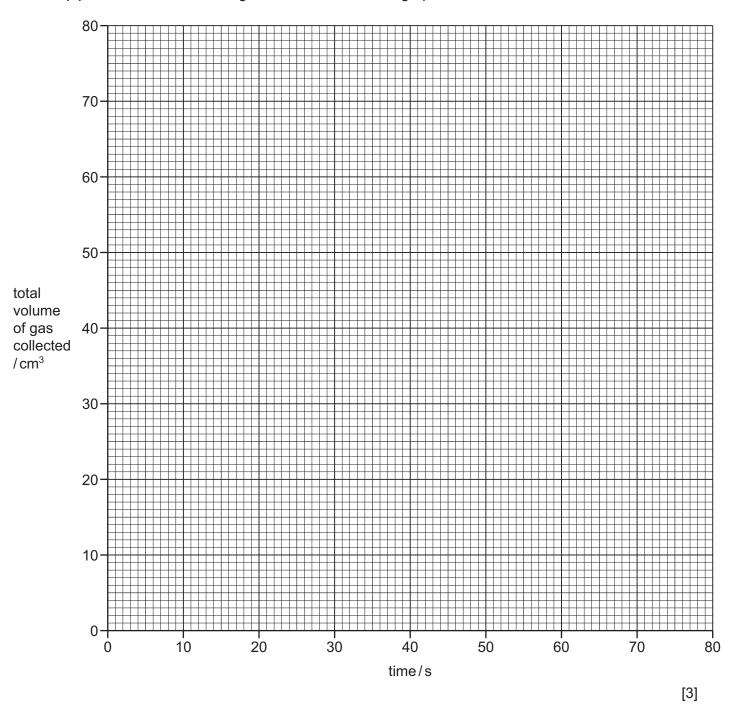


50 cm³ of dilute nitric acid was added to the conical flask. The contents of the test-tube were released, allowing the solid magnesium carbonate to mix with the dilute nitric acid. A stop-clock was started and the volume of gas collected in the inverted measuring cylinder was measured every 10 seconds for 80 seconds.

(a) Use the inverted measuring cylinder diagrams to record the volume of gas collected in the table

time/s	inverted measuring cylinder diagram	total volume of gas collected/cm³
0	9 9 1 01	
10		
20	07- 97- 99-	
30	99 = 99 = 90 =	
40	09-= 	
50	02 = 92 = 08 =	
60	02 = 92 = 08 = 08 =	
70	02 = 92 = 08 =	
80	02 = 92 = 08 = 08 =	





(c) (i) One of the points is anomalous.

Circle this point on your graph. [1]

(ii) From your graph, deduce the time taken to collect 40 cm³ of gas. Show clearly on the grid how you worked out your answer.

			_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	_	S		Ľ	2	1
		•																											_	,	- 14	_	ı

(d)		te one possible source of error in this experiment. Suggest one improvement to reduce this rce of error.
	sou	rce of error
	imp	rovement[2]
(e)	The	average rate of the reaction can be calculated using the equation shown.
		average rate of reaction = $\frac{\text{volume of gas collected/cm}^3}{\text{time taken/s}}$
	(i)	Calculate the volume of gas collected between 10 seconds and 30 seconds.
	(ii)	Calculate the average rate of reaction between 10 seconds and 30 seconds. Include the unit in your answer.
		average rate of reaction = unit =
		[2]
(f)	The	student calculated that the total volume of gas collected in this reaction would be 85 cm ³ .
	Sug	igest and explain why the actual volume of gas collected was different from 85 cm ³ .
		[2]
(g)		tch on the grid the graph you would expect if the experiment were repeated at a lower perature. Label this graph as L . [2]
		[Total: 17]

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3 Two substances, solid **A** and solid **B**, were analysed. Solid **A** was zinc nitrate. Tests were done on the substances.

tests on solid A

Complete the expected observations.

Solid **A** was added to distilled water and the mixture shaken to dissolve solid **A** and produce solution **A**. Solution **A** was divided into three equal portions in three test-tubes.

(a)	(i)	A few drops of aqueous sodium hydroxide were added to the first portion of solution A .	
		observations	[2]
	(ii)	An excess of aqueous sodium hydroxide was then added to this mixture.	
		observations	[1]
(b)	(i)	A few drops of aqueous ammonia were added to the second portion of solution A .	
		observations	[1]
	(ii)	An excess of aqueous ammonia was then added to this mixture.	
		observations	[1]
(c)		minium foil and aqueous sodium hydroxide were added to the third portion of solution e mixture was heated and the gas produced was tested with litmus paper.	A.
	obs	ervations	
			2

tests on solid B

Some of the tests and observations are shown.

tests on solid B	observations
test 1	
A flame test was done on solid B .	red flame
test 2	
Solid B was dissolved in water.	
Dilute nitric acid and aqueous silver nitrate were added to the solution.	yellow precipitate formed

(d)	Identify solid B .	
		[2]

lodine dissolves in two different solvents: ethanol and hexane.

Plan an experiment to find out in which solvent iodine is the most soluble at room temperature.
You are provided with iodine, the two solvents and common laboratory apparatus.

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